

COVALENT NOMENCLATURE



Thank you for downloading the **Covalent Nomenclature Interactive Graphic Organizers**. I appreciate your business. If I may ask, please take a moment of your time to provide feedback *after you have used the product* to help me gauge your level of satisfaction in regard to classroom usage. Specific feedback helps me to modify and/or generate products. Should you encounter technical issues, please contact me via email or the Q&A tab on my page prior to feedback as this may not be a direct result of the product itself. Additionally, please check back periodically as I usually make updates to the product either based on questions submitted, your feedback, or simply wanting to add to the product itself.

Once again, thank you for your time and business. Have a school great year!

—James Lyas (Bond with James)—

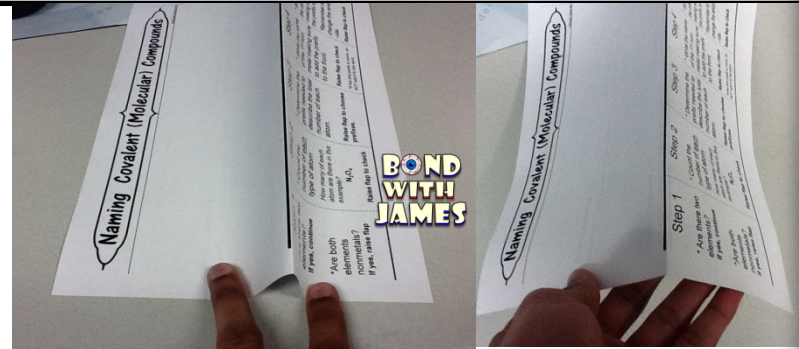
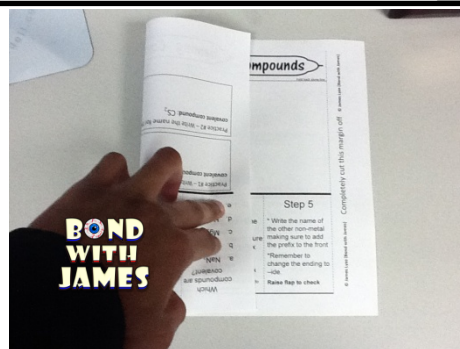
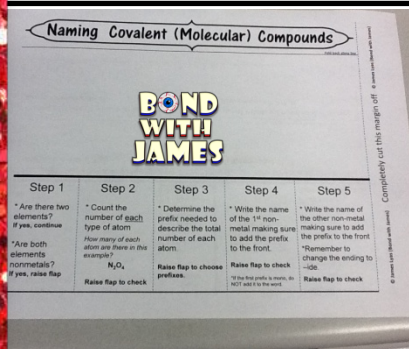
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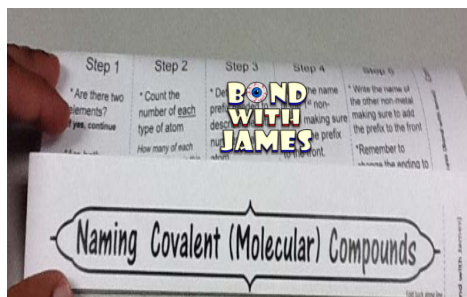
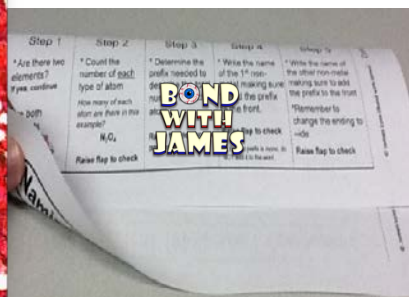
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INTERACTIVE GRAPHIC ORGANIZER DIRECTIONS & SUGGESTIONS FOR USE

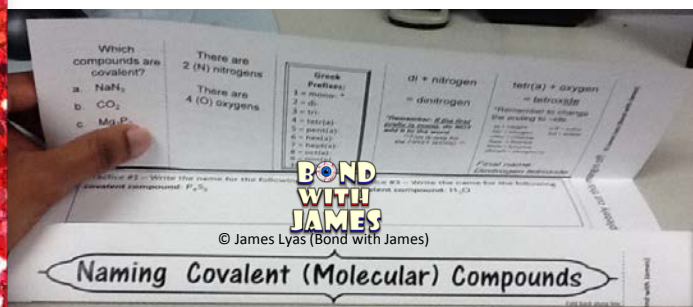
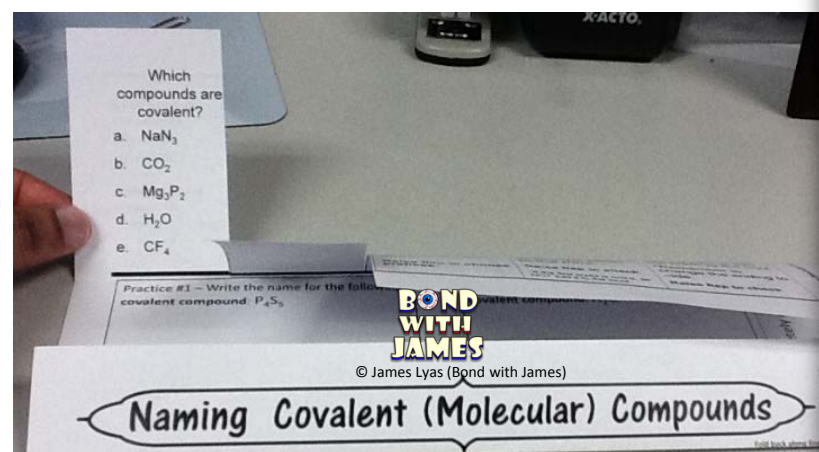


Step 1: Make sure to print so that the tab sections on both sides are behind one another (as shown here)

Step 2: On the side that has the organizer title, fold paper along the solid bold line.



Step 3: Fold the organizer title graphic back (as indicated on thin black line)



Step 4: Finally, cut on dotted lines between each tab up to the bold line.

Students can tuck folds behind Title tab.

****Problem 4 is wrong on PURPOSE. If students truly understand, they will catch this and recognize that they cannot apply covalent nomenclature to ionic compounds.**

Naming Covalent (Molecular) Compounds

Fold back along line

Fold along line

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Step 1

* Are there two elements?
If yes, continue

* Are both elements nonmetals?
If yes, raise flap

Step 2

* Count the number of each type of atom

How many of each atom are there in this example?



Raise flap to check

Step 3

* Determine the prefix needed to describe the total number of each atom.

Raise flap to choose prefixes.

Step 4

* Write the name of the 1st non-metal making sure to add the prefix to the front.

Raise flap to check

**If the first prefix is mono, do NOT add it to the word.*

Step 5

* Write the name of the other non-metal making sure to add the prefix to the front

* Remember to change the ending to -ide.

Raise flap to check

Completely cut off margins with dotted lines

Which compounds are covalent?

- a. NaN_3
- b. CO_2
- c. Mg_3P_2
- d. H_2O
- e. CF_4

There are 2 (N) nitrogens

There are 4 (O) oxygens

Greek Prefixes:

- 1 = mono- *
- 2 = di-
- 3 = tri-
- 4 = tetr(a)-
- 5 = pent(a)-
- 6 = hex(a)-
- 7 = hept(a)-
- 8 = oct(a)-
- 9 = non(a)-
- 10 = dec(a)-

di + nitrogen
= dinitrogen

****Remember: If the first prefix is mono, do NOT add it to the word.***

*****This is only for the FIRST WORD.*****

tetr(a) + oxygen
= tetroxide

***Remember to change the ending to -ide.**

ox = oxygen	sulf = sulfur
nitr = nitrogen	iod = iodine
chlor = chlorine	
fluor- = fluorine	
brom = bromine	
phosph = phosphorus	

***Final name:
Dinitrogen tetroxide***

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Practice #1 – Write the name for the following covalent compound: P_4S_5

Practice #3 – Write the name for the following covalent compound: H_2O

Practice #2 – Write the name for the following covalent compound: CS_2

Practice #4 – Write the name for the following covalent compound: Mg_3N_2